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~~not do " long division " to try to get exact values. Remember it is a MC test, use the answers • Mark which questions you would like to " go over " when we get to school in September. 1. Balance the following equation: ___ NH₃ + ___ O₂ → ___ NO₂ + ___ H₂O The balanced equation shows that 1.00 mole of NH₃ requires ___ mole(s) of O₂ a. 0.57 b. 1.25 c. 1.33~~

~~Practice Test Ch 3 Stoichiometry Name Per~~

~~fewer steps are required to solve stoichiometry problems when the reactant is given in moles and the product is sought in moles which of the following mathematical expressions correctly states the relationship among percentage yield, actual yield, and theoretical yield %yield= (act. yield/theo. yield) × 100~~

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~~Chapter 9 Stoichiometry Test Answers~~

Practice Problems (Chapter 5): Stoichiometry CHEM 30A Part I: Using the conversion factors in your tool box g A mol A mol A 1. How many moles CH₃OH are in 14.8 g CH₃OH? 2. What is the mass in grams of 1.5 x 10¹⁶ atoms S? 3. How many molecules of CO₂ are in 12.0 g CO₂? 2 4. What is the mass in grams of 1 atom of Au? KEY Tool Box: To ...

~~Practice Problems (Chapter 5): Stoichiometry~~

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~~Chemistry Matter And Change Chapter 12 Stoichiometry Study ...~~

Stoichiometry Chapter 3! Stoichiometry: Calculations with Chemical Formulas and Equations. Stoichiometry Anatomy of a Chemical Equation CH₄ (g) + 2O₂ (g) CO₂ (g) + 2H₂O (g) Stoichiometry Anatomy of a Chemical Equation Reactants appear on the left side of the equation. CH₄ (g) + 2O₂ (g) CO₂ (g) + 2H₂ ...

~~Chapter 3 Stoichiometry—Chemistry~~

Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemi-cal reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

~~Chapter 11: Stoichiometry~~

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~~Chapter 03—Stoichiometry~~

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~~Sample Exam Questions~~

Answer outline and marking scheme for question: 2. Molar mass of HgO = 201 + 16 = 217 g mol⁻¹. 1.08g of HgO will contain 1.08 / 217 mols = 0.005mol From the equation, 1 mole of O₂(g) is produced from 2 moles of HgO This means that 0.005 mol of HgO will produce 0.005 / 2 mol of O₂ = 0.0025 mol 0.0025 mol of O₂ will occupy 0.0025 x 24dm³ = 0.06dm³

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