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

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chemistry where

T_0 is the

freezing point
of the solvent.

If BC in Fig.

1st-1 is taken
to be a straight
line, the slope

dp_A/dT may be
replaced by

CB/BD, i.e., by

$$(p_A - p_s) / \Delta T,$$

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so that $\ln \frac{p}{p_0} = -\frac{H_p}{RT} + \frac{A}{T} + B$
Eq. 1st.4

The slope of the vapor pressure curve of the solid solvent is also given by another form of the

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The dilution formula allows you to determine the molarity of a solution after it has been diluted. The formula is: $M_1V_1 = M_2V_2$

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$M_1 V_1 = M_2 V_2$ Where M_1 and V_1 are the molarity and volume before a dilution, and M_2 and V_2 are the molarity and volume after a dilution.

Consider this example from the 2006 NECTA:

"Hydrochloric acid solution

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Figure 1.1

Packing models
of (a) side-by-
side layer, (b)
hexagon layer,
(c) simple cubic
layer, (d) body
centered cubic.
Determine the
maximum amount

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of balls that
can touch
another one ball
(center- ball)
in the same
layer for each
arrangement .

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