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11. molar mass = mass/moles; mass = 45.0 g; find moles; ? T
b = $90.2^{\circ}\text{C} - 80.1^{\circ}\text{C} = 10.1^{\circ}\text{C}$ $\left(\text{moles} = \frac{\Delta T_f}{2.53} \times \text{kg benzene} = \frac{10.1^{\circ}\text{C}}{2.53} \times 0.0750 = 0.300 \text{ moles}\right)$

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Ans: Primary alcohols: (i), (ii), (iii) Secondary alcohols: (iv), (v) Tertiary alcohols: (vi) 11.2. Identify aliphatic alcohols in the above examples. Ans: (ii) and (iv) i.e. $\text{H}_2\text{C}=\text{CH}-\text{CH}_2\text{OH}$

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Question 11. What would be the reactant and reagent used to obtain 2, 4-dimethyl pentan-3-ol? (a) Propanal and propyl magnesium bromide (b) 3-methylbutanal and 2-methyl magnesium iodide (c) 2-dimethylpropanone and methyl magnesium iodide (d) 2-methylpropanal and isopropyl magnesium iodide. Answer: (d) 2-methylpropanal and isopropyl magnesium iodide

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Answer: The basic theme of organisation of elements in the periodic table is to simplify and systematize the study of the

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properties of all the elements and millions of their compounds. This has made the study simple because the properties of elements are now studied in form of groups rather than individually.

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