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$$[K] = [^{\circ}C] + 273.15$$

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Another absolute temperature scale – the Rankine scale – can be used to convert between SI and English systems:  
 $[\text{°R}] = 1.8 [\text{K}]$   
 $[\text{°R}] = [\text{°F}] + 459.67$  Of course, one can then write that  
 $[\text{°F}] = 1.8 [\text{°C}] + 32$

### 1.2.3 PRESSURE

Pressure,  $P$ , is also an intensive property.

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Many of the definitions  
below are also used in  
the thermodynamics of  
chemical reactions.

General basic quantities.

Quantity (Common  
Name/s) (Common)

Symbol/s SI Units

Dimension Number of  
molecules N:

dimensionless

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dimensionless Number  
of moles  $n$ : mol ... C V  
 $= ? / ? J K ?^1 [M][L ...$

Table of thermodynamic  
equations - Wikipedia

In thermodynamics and  
chemical engineering,  
the vapor–liquid  
equilibrium (VLE)  
describes the  
distribution of a  
chemical species  
between the vapor phase



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and a liquid phase.. The concentration of a vapor in contact with its liquid, especially at equilibrium, is often expressed in terms of vapor pressure, which will be a partial pressure (a part of the total gas pressure) if any other gas(es ...

Vapor–liquid  
equilibrium - Wikipedia

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thermodynamics... Thermodynamics sounds intimidating, and it can be. However, if you hone in on the most important thermodynamic formulas and equations, get comfortable converting from one unit of physical measurement to another, and become familiar with the physical

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constants related to thermodynamics, you'll be at the head of the class.

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Krypton (from  
Greek:??????? kryptos  
"the hidden one") is a  
chemical element with  
symbol Kr and atomic  
number 36. It is a

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member of group 18  
(noble gases) elements.  
It is a member of group  
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