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1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage

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yield. 2. 6.0

mol of N_2 are
mixed with 12.0
mol of H_2

according to the
following
equation: $N_2(g)$
 $3H_2(g) \dots$

mc06se cFMsr i-
vi -

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the answer on

the line to the

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your work in the

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45 mol The

following

equation

represents a

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laboratory
preparation for
oxygen gas:



How many moles
of O_2 form if
30 mol of

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provide an
efficient and
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Calculate the
molar mass of
the following,
and then write
the molar mass

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in ratio form.

KNO_3 . Na_2SO_4 .

$\text{Ca}(\text{OH})_2$

$(\text{NH}_4)_2\text{SO}_3$.

CuSO_4 . Moles and

Mass. Determine

the ...

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equation allows
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one to the
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mole ratio of
any two
substances in
the reaction.
The coefficients
in a chemical
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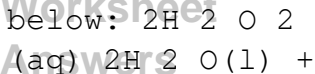
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Answers

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Answers will vary. An example is included



$\text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide.

Microscopic: Two molecules of hydrogen

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peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

Macroscopic: Two moles of hydrogen peroxide (present in ...

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the answer on
the line to the
left. Show all.

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Equations 11.1

11.2 $U = mgh$ $E = K$

$U = \frac{1}{2}mv^2$ Chapter

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space provided 1
88% The actual
yield of a
reaction is 22 g
and the

theoretical
yield is 25 g
Calculate the
percentage yield

2 60 mol of N₂
are mixed with
120 mol of H

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