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Equation 7.3.1 and Equation 7.3.2 are balanced chemical equations. What is different on each side of the equation is how the atoms are arranged to make molecules or ions. A chemical reaction represents a change in the distribution of atoms but not in the number of atoms. In this reaction, and in most chemical reactions, bonds are

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broken in the reactants (here, Cr–O and N–H bonds), and new bonds are formed to create the products (here, O–H and N–N bonds).

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4 a S in H₂SO₃ b

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Chapter 7 3 Answers Chemical CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x

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g/mol. False b. Samples

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following

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questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples of equal numbers of moles of two different chemicals

7 Chemical Formulas and Chemical Compounds

Chapter 7 3 Answers Chemical Formulas And Chemical Compounds includes public domain books and promotional books that legal copyright holders wanted to give away for free. Chapter 7 3 Answers Chemical As shown in Figure 7.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the Page 4/30

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ethanol ($\text{CH}_3\text{CH}_2\text{OH}$) Answer: Ba, +2; F, ?1; C, 0; H, +1; O, ?2; K, +1; Cr, +6; O, ?2; Cs, +1; O, ? $\frac{1}{2}$; C, ?3; H, +1; C, ?1; H, +1; O, ?2; H, +1

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Chapter 7.5: Types of Chemical Reactions - Chemistry ...

Answer. (a) baking soda (sodium bicarbonate) plus vinegar (acetic acid) yields carbon dioxide plus water plus sodium ion plus acetate ion. Vinegar (Acetic acid) + Baking soda (Sodium hydrogencarbonate) -> Carbon dioxide + other substances.
$$\text{NaHCO}_3 + \text{CH}_3\text{COOH} \rightarrow \text{NaC}_2\text{H}_3\text{O}_2 + \text{H}_2\text{O} + \text{CO}_2$$

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a) b) 3. c) Writing and Balancing Chemical Equations. 5. b) 7. a)

$\text{Zn(s)} + 2\text{HCl(aq)} \rightarrow \text{ZnCl}_2\text{(aq)} + \text{H}_2\text{(g)}$ b) $\text{CH}_4\text{(g)} + 2\text{O}_2\text{(g)} \rightarrow$

$\text{CO}_2\text{(g)} + 2\text{H}_2\text{O(l)}$ c) $\text{Ca(OH)}_2\text{(aq)} + \text{CO}_2\text{(g)} \rightarrow \text{CaCO}_3\text{(s)} +$

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H₂O(l) d) 2Fe₂O₃ ...

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Iron and sulfur react to form Iron (III) sulfide 2. Phosphorus and oxygen combine to form diphosphorus pentoxide 3. Water decomposes to form hydrogen and oxygen 4. Sodium reacts with water to form sodium hydroxide and hydrogen 5. Iron metal reacts with copper (II) nitrate to form iron (II) nitrate and copper metal 6.

Class Worksheet-2 Chapter 7 Name More Practice Pre ...

Balancing Chemical Equations Chapter 7 Balancing Chemical

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Equations Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightleftharpoons 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightleftharpoons 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightleftharpoons 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightleftharpoons 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightleftharpoons$
2

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