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Atmospheric pressure decreases higher up in the atmosphere.

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Answer. Molar mass = 151.1 g/mol

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Potassium: $(39.10 / 101.1) \times 100 = 38.67\%$
Nitrogen: $(14.01 / 101.1) \times 100 = 13.86\%$
Oxygen: $(48.00 / 101.1) \times 100 = 47.48\%$. Molar mass = 164.1 g/mol

Potassium: $(39.10 / 101.1) \times 100 = 38.67\%$
Nitrogen: $(14.01 / 101.1) \times 100 = 13.86\%$
Oxygen: $(48.00 / 101.1) \times 100 = 47.48\%$.

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11-1 At the same temperature and
pressure, balloons of equal volume
have equal numbers of molecules,
regardless of which gas they contain.

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Hydrogen molecule 1 mol H_2 at STP
= 22.4 L Oxygen molecule 1 mol O_2
at STP = 22.4 L Carbon dioxide
molecule 1 mol CO_2 at STP = 22.4 L
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nitrogen gas, 21% oxygen gas, and

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1% other gases by volume.

Atmospheric pressure decreases
higher up in the atmosphere.

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