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11. molar mass = mass/moles; mass = 45.0 g; find moles; ? T
b = $90.2^{\circ}\text{C} - 80.1^{\circ}\text{C} = 10.1^{\circ}\text{C}$ $\left(\text{moles} = \frac{\Delta}{\text{molar mass}}\right)$

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$$T_{f\{2.53\}} \times \text{kg} \text{ benzene} = \frac{10.1 \text{ g}}{2.53 \text{ g/mol}} = 0.300 \text{ moles}$$

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Answer/Explanation. Answer: Explanation: (i) (d) (ii) (e) (iii) (b) (iv) (a) In cleavage of ethers, higher alcohol and lower alkyl halide is formed except in tert. butyl methyl ether, in which $(\text{CH}_3)_3\text{C}^+$ is more stable so $(\text{CH}_3)_3\text{C-I}$ and CH_3OH is formed. Anisole with HI gives phenol and CH_3I because phenoxide ion is stabilised by resonance.

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