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structure of the

atom Although

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the ancient  
Greeks had many  
creative ideas

about matter,

our current

understanding is

based upon John

Dalton's atomic

theory,

published in

1803.

Essentially,

matter is made

up of atoms,

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Structure  
Chapter 4  
which are  
indivisible and  
indestructible.

Chapters 4:

Atomic Structure

- ANNE SCHMIDT

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preview: Chapter

4 "Atomic

Structure"

Pearson

Chemistry Book

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High School 1

OBJECTIVES:

1. Defining the atom. 2.

Recognize the change of the atomic structure through history.

3. Differentiate between the properties of subatomic particles.

4. Describe the

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Structure  
Characteristics  
of the atom. 2

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I. Defining the  
Atom Warm up ( 1  
min ) What is an  
atom? 3 Living  
things are made  
up of Cells ...

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t answers ...

4.05 Atomic

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Structure, Mass

Number, Isotopes  
and Isobars 4.05

Atomic Number,

Mass number,

Isotopes and

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Plasma Membrane  
5.02 Plasma  
Membrane ...

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Structure of

Atom ...

1. All elements  
are composed of

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Structure  
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tiny invisible particles

(atoms) 2. All atoms of the same element are identical. 3. Atoms of different elements can physically combine together or chemically combine in simple who

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number ratios to  
form compounds.

4. Chemical

reactions occur

when atoms are

separated,

joined, or

rearranged.

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theory. 4 part

theory about

atoms. 1) All

elements are

composed of

\_\_\_\_\_.

atoms. Atoms can

combine both

\_\_\_\_\_ and

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## Atomic

## Structure

physically &  
chemically. 2)

atoms in a given  
element are

\_\_\_\_\_, but

different from  
those in other  
elements.

identical. 3)

atoms are

neither \_\_\_\_\_

nor \_\_\_\_\_.

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Concepts covered

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Atomic Structure

are Fundamental

Subatomic

Particles

Present in an

Atom: Electrons,

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Protons,  
Neutrons,  
Chapter 4  
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Extra Nuclear

Parts, Concept  
of Atomic

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Atoms, element.

Atoms are

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indivisible and  
indestructible.

Nucleus is

small, dense,

and positively

charged (is the

explin... Whose

Experiment

proved atoms are

neutral,

electrons have

very... .

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The nucleus has two fewer protons, so the atomic number decreases by two. The nucleus also has two fewer neutrons, so the mass number decreases by four. A new element is formed.

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Structure for

sulfur is 16.

How many

electrons are

there in an atom

of sulfur-34?

answer choices .

16. 18. 34. 50.

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Worksheet Answer

Key Measuring

Atomic Mass

Instead of

grams, the unit

we use is the

Atomic Mass Unit

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(amu) It is

defined as one-twelfth the mass of a carbon-12 atom. Carbon-12 chosen because of its isotope purity.

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nucleus of the atom contains the neutrons and positively charged protons. The nucleus is surrounded by as many orbital electrons (each with a charge of  $-1$ ) as there are protons (each with a charge of  $+1$ ) in



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Chapter 4

the nucleus. The three sub-atomic particles differ from each other in two important ways.

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